JEE Advanced (2024)

PAPER 1

General Instructions:

SECTION 1 (Maximum Marks: 12)

- This section contains **FOUR (04)** questions.
- Each question has **FOUR** options (A), (B), (C) and (D). **ONLY ONE** of these four options is the correct answer.
- For each question, choose the option corresponding to the correct answer.
- Answer to each question will be evaluated according to the following marking scheme:

Full Marks	:	+3 If **ONLY** the correct option is chosen;
Zero Marks	:	0 If none of the options is chosen (i.e., the question is unanswered);
Negative Marks	:	−1 in all other cases.

1. A closed vessel contains 10 g of an ideal gas X at 300 K, which exerts 2 atm pressure. At the same temperature, 80 g of another ideal gas Y is added to it and the pressure becomes 6 atm. The ratio of root mean square velocities of X and Y at 300 K is

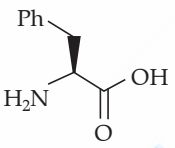
(A) $2\sqrt{2} : \sqrt{3}$ **(B)** $2\sqrt{2} : 1$ **(C)** 1 : 2 **(D)** 2 : 1

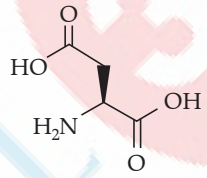
2. At room temperature, disproportionation of an aqueous solution of *in situ* generated nitrous acid (HNO_2) gives the species

(A) H_3O^+, NO_3^- and NO **(B)** H_3O^+, NO_3^- and NO_2
(C) H_3O^+, NO^- and NO_2 **(D)** H_3O^+, NO_3^- and N_2O

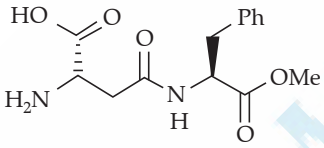
3. Aspartame, an artificial sweetener, is a dipeptide aspartyl phenylalanine methyl ester. The structure of aspartame is

Structures of phenylalanine and aspartic acid are given below.

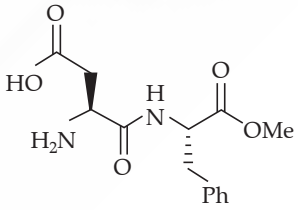




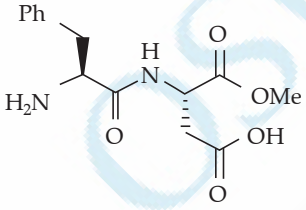
(A)



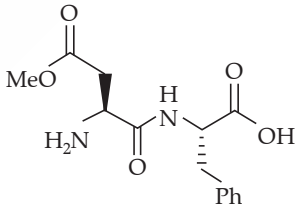
(B)



(C)



(D)



4. Among the following options, select the option in which each complex in **Set-I** shows geometrical isomerism and the two complexes in **Set-II** are ionisation isomers of each other.

[en = $H_2NCH_2CH_2NH_2$]

(A) **Set-I:** $[Ni(CO)_4]$ and $[PdCl_2(PPh_3)_2]$
Set-II: $[Co(NH_3)_5Cl]SO_4$ and $[Co(NH_3)_5(SO_4)]Cl$

(B) **Set-I:** $[Co(en)(NH_3)_2Cl_2]$ and $[PdCl_2(PPh_3)_2]$
Set-II: $[Co(NH_3)_6][Cr(CN)_6]$ and $[Cr(NH_3)_6][Co(CN)_6]$

(C) **Set-I:** $[Co(NH_3)_3(NO_2)_3]$ and $[Co(en)_2Cl_2]$
Set-II: $[Co(NH_3)_5Cl]SO_4$ and $[Co(NH_3)_5(SO_4)]Cl$

(D) **Set-I:** $[Cr(NH_3)_5Cl]Cl_2$ and $[Co(en)(NH_3)_2Cl_2]$
Set-II: $[Cr(H_2O)_6]Cl_3$ and $[Cr(H_2O)_5Cl]Cl_2 \cdot H_2O$

General Instructions:

SECTION 2 (Maximum Marks: 12)

- This section contains **THREE (03)** questions.
- Each question has **FOUR** options (A), (B), (C) and (D). **ONE OR MORE THAN ONE** of these four option(s) is (are) correct answer(s).
- For each question, choose the option(s) corresponding to (all) the correct answer(s).
- Answer to each question will be evaluated according to the following marking scheme:

Full Marks	:	+4 ONLY if (all) the correct option(s) is(are) chosen;
Partial Marks	:	+3 If all the four options are correct but ONLY three options are chosen;
Partial Marks	:	+2 If three or more options are correct but ONLY two options are chosen, both of which are correct;
Partial Marks	:	+1 If two or more options are correct but ONLY one option is chosen and it is a correct option;
Zero Marks	:	0 If unanswered;
Negative Marks	:	−2 in all other cases.

- For example, in a question, if (A), (B) and (D) are the ONLY three options corresponding to correct answers, then
  choosing ONLY (A), (B) and (D) will get +4 marks;
  choosing ONLY (A) and (B) will get +2 marks;
  choosing ONLY (A) and (D) will get +2 marks;
  choosing ONLY (B) and (D) will get +2 marks;
  choosing ONLY (A) will get +1 mark;
  choosing ONLY (B) will get +1 mark;
  choosing ONLY (D) will get +1 mark;
  choosing no option(s) (i.e., the question is unanswered) will get 0 marks and
  choosing any other option(s) will get −2 marks.

5. Among the following, the correct statement(s) for electrons in an atom is(are)

(A) Uncertainty principle rules out the existence of definite paths for electrons.

(B) The energy of an electron in 2*s* orbital of an atom is lower than the energy of an electron that is infinitely far away from the nucleus.

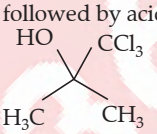
(C) According to Bohr's model, the most negative energy value for an electron is given by $n = 1$, which corresponds to the most stable orbit.

(D) According to Bohr's model, the magnitude of velocity of electrons increases with increase in values of n.

6. Reaction of iso-propylbenzene with O_2 followed by the treatment with H_3O^+ forms phenol and a by-product **P**. Reaction of **P** with 3 equivalents of Cl_2 gives compound *Q*. Treatment of **Q** with $Ca(OH)_2$ produces compound *R* and calcium salt **S**.

The correct statement(s) regarding **P**, **Q**, **R** and **S** is(are)

(A) Reaction of **P** with **R** in the presence of KOH followed by acidification gives



(B) Reaction of **R** with O_2 in the presence of light gives phosgene gas.

(C) **Q** reacts with aqueous NaOH to produce Cl_3CCH_2OH and $Cl_3CCOONa$

(D) **S** on heating gives **P**

7. The option(s) in which at least three molecules follow Octet Rule is(are)

(A) CO_2, C_2H_4, NO and HCl

(B) NO_2, O_3, HCl and H_2SO_4

(C) BCl_3, NO, NO_2 and H_2SO_4

(D) CO_2, BCl_3, O_3 and C_2H_4

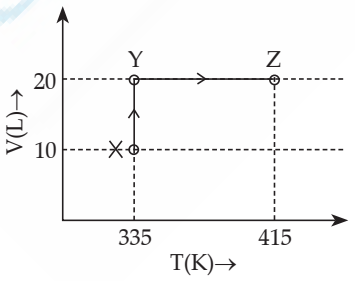
General Instructions:

SECTION 3 (Maximum Marks: 24)

- This section contains **SIX (06)** questions.
- The answer to each question is a **NON-NEGATIVE INTEGER**.
- For each question, enter the correct integer corresponding to the answer using the mouse and the onscreen virtual numeric keypad in the place designated to enter the answer.
- Answer to each question will be evaluated according to the following marking scheme:

Full Marks	:	+4 If **ONLY** the correct integer is entered;
Zero Marks	:	0 in all other cases.

8. Consider the following volume−temperature (V−T) diagram for the expansion of 5 moles of an ideal monoatomic gas.



Considering only P-V work is involved, the total change in enthalpy (in Joule) for the transformation of state in the sequence **X** → **Y** → **Z** is _____.

[Use the given data: Molar heat capacity of the gas for the given temperature range, $C_{V,m} = 12\ J\ K^{-1}\ mol^{-1}$ and gas constant, $R = 8.3\ J\ K^{-1}\ mol^{-1}$]

9. Consider the following reaction,

$2H_2(g) + 2NO(g) \rightarrow 2N_2(g) + 2H_2O(g)$

which follows the mechanism given below:

$2NO(g) \underset{k_{-1}}{\overset{k_1}{\rightleftharpoons}} N_2O_2(g)$ (fast equilibrium)

$N_2O_2(g) + H_2(g) \xrightarrow{k_2} N_2O(g) + H_2O(g)$ (slow reaction)

$N_2O(g) + H_2(g) \xrightarrow{k_3} N_2(g) + H_2O(g)$ (fast reaction)

The order of the reaction is _____.

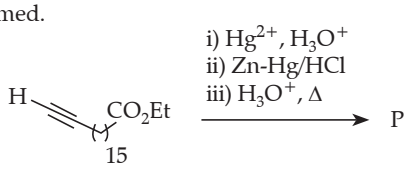
10. Complete reaction of acetaldehyde with excess formaldehyde, upon heating with conc. NaOH solution, gives **P** and **Q**. Compound **P** does not give Tollens' test, whereas **Q** on acidification gives positive Tollens' test. Treatment of **P** with excess cyclohexanone in the presence of catalytic amount of p-toluenesulfonic acid (PTSA) gives product **R**.

Sum of the number of methylene groups (-CH_2-) and oxygen atoms in **R** is _____.

11. Among $V(CO)_6$, $Cr(CO)_5$, $Cu(CO)_3$, $Mn(CO)_5$, $Fe(CO)_5$, $[Co(CO)_3]^{3-}$, $[Cr(CO)_4]^{4-}$ and $Ir(CO)_3$, the total number of species isoelectronic with $Ni(CO)_4$ is _____.

[Given, atomic number: V = 23, Cr = 24, Mn = 25, Fe = 26, Co = 27, Ni = 28, Cu = 29, Ir = 77]

12. In the following reaction sequence, the major product **P** is formed.



Glycerol reacts completely with excess **P** in the presence of an acid catalyst to form **Q**. Reaction of Q with excess NaOH followed by the treatment with $CaCl_2$ yields Ca-soap **R**, quantitatively.

Starting with one mole of **Q**, the amount of **R** produced in gram is _____.

[Given, atomic weight: H = 1, C = 12, N = 14, O = 16, Na = 23, Cl = 35, Ca = 40]

13. Among the following complexes, the total number of diamagnetic species is _____.

$[Mn(NH_3)_6]^{3+}$, $[MnCl_6]^{3-}$, $[FeF_6]^{3-}$, $[CoF_6]^{3-}$, $[Fe(NH_3)_6]^{3+}$, and $[Co(en)_3]^{3+}$

[Given, atomic number: Mn = 25, Fe = 26, Co = 27; en = $H_2NCH_2CH_2NH_2$]

General Instructions:

SECTION 4 (Maximum Marks: 12)

- This section contains **FOUR (04)** paragraphs.
- Each set has **ONE** Multiple Choice Question.
- Each set has **TWO** lists: **List-I** and **List-II.**
- **List-I** has **Four** entries (P), (Q), (R) and (S) and **List-II** has **FIVE** entries (1), (2), (3), (4) and (5).
- **FOUR** options are given in each Multiple Choice Question based on **List-I** and **List-II** and **ONLY ONE** of these four options satisfies the condition asked in the Multiple Choice Question.
- Answer to each question will be evaluated according to the following marking scheme:

Full Marks	:	+3	**ONLY** if the option corresponding to the correct combination is chosen.
Zero Marks	:	0	If none of the options is chosen (i.e the question is unanswered);
Negative Marks	:	–1	in all other cases.

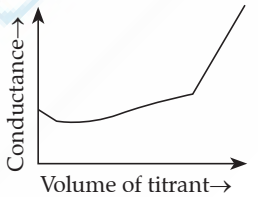
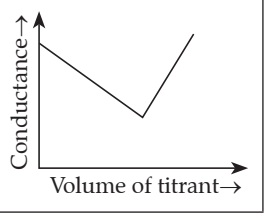
14. In a conductometric titration, small volume of titrant of higher concentration is added stepwise to a larger volume of titrate of much lower concentration, and the conductance is measured after each addition.

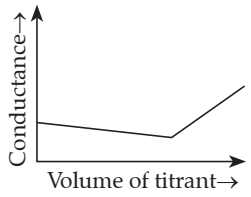
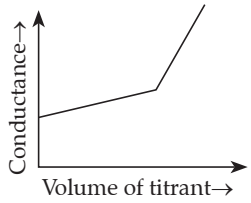
The limiting ionic conductivity (Λ_0) values (in mS m^2 mol^{-1}) for different ions in aqueous solutions are given below:

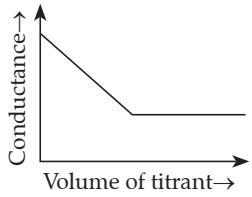
Ions	Ag^+	K^+	Na^+	H^+	NO_3^-	Cl^-	SO_4^{2-}	OH^-	CH_3COO^-
Λ_0	6.2	7.4	5.0	35.0	7.2	7.6	16.0	19.9	4.1

For different combinations of titrates and titrants given in **List-I**, the graphs of 'conductance' versus 'volume of titrant' are given in **List-II**.

Match each entry in **List-I** with the appropriate entry in **List-II** and choose the correct option.

List-I	List-II
(P) Titrate: KCl Titrant: $AgNO_3$	(1) 
(Q) Titrate: $AgNO_3$ Titrant: KCl	(2) 

(R) Titrate: NaOH Titrant: HCl	(3) 
(S) Titrate: NaOH Titrant: CH_3COOH	(4) 

	(5) 

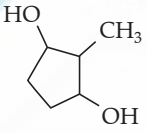
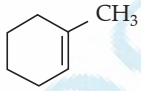
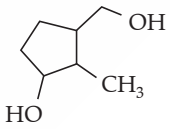
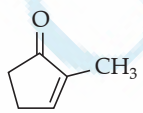
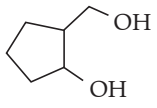
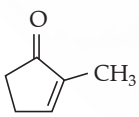
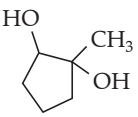
(A) P → 4; Q → 3; R → 2; S → 5
(B) P → 2; Q → 4; R → 3; S → 1
(C) **P → 3; Q → 4; R → 2; S → 5**
(D) **P → 4; Q → 3; R → 2; S → 1**

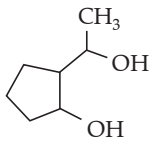
15. Based on VSEPR model, match the xenon compounds given in **List-I** with the corresponding geometries and the number of lone pairs on xenon given in **List-II** and choose the correct option.

List-I	List-II
(P) XeF_2	(1) Trigonal bipyramidal and two lone pair of electrons
(Q) XeF_4	(2) Tetrahedral and one lone pair of electrons
(R) XeO_3	(3) Octahedral and two lone pair of electrons
(S) XeO_3F_2	(4) Trigonal bipyramidal and no lone pair of electrons
	(5) Trigonal bipyramidal and three lone pair of electrons

(A) P → 5; Q → 2; R → 3; S → 1
(B) P → 5; Q → 3; R → 2; S → 4
(C) P → 4; Q → 3; R → 2; S → 1
(D) P → 4; Q → 2; R → 5; S → 3

16. **List-I** contains various reaction sequences and **List-II** contains the possible products.
Match each entry in **List-I** with the appropriate entry in **List-II** and choose the correct option.

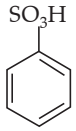
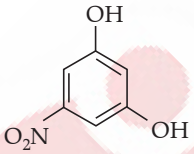
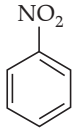
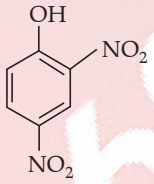
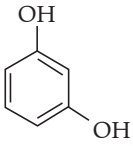
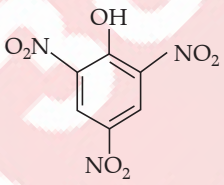
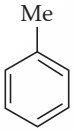
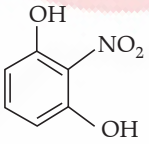
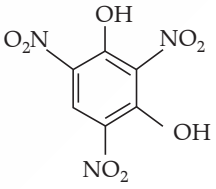
List-I	List-II
(P)  (i) O_3, Zn; (ii) aq. NaOH, Δ; (iii) ethylene glycol, PTSA; (iv) (a) BH_3, (b) H_2O_2, NaOH; (v) H_3O^+; (vi) $NaBH_4$	(1) 
(Q)  (i) O_3, Zn; (ii) aq. NaOH, Δ; (iii) ethylene glycol, PTSA; (iv) (a) BH_3, (b) H_2O_2, NaOH; (v) H_3O^+; (vi) $NaBH_4$	(2) 
(R)  (i) ethylene glycol, PTSA; (ii) (a) $Hg(OAc)_2H_2O$, (b) $NaBH_4$; (iii) H_3O^+; (iv) $NaBH_4$	(3) 
(S)  (i) ethylene glycol, PTSA; (ii) a) BH_3, b) H_2O_2, NaOH; (iii) H_3O^+; (iv) $NaBH_4$	(4) 

	(5) 

(A) P → 3; Q → 5; R → 4; S → 1
(B) P → 3; Q → 2; R → 4; S → 1
(C) P → 3; Q → 5; R → 1; S → 4
(D) P → 5; Q → 2; R → 4; S → 1

17. **List-I** contains various reaction sequences and **List-II** contains different phenolic compounds.

Match each entry in **List-I** with the appropriate entry in **List-II** and choose the correct option.

List-I	List-II
(P)  (i) molten NaOH, H_3O^+ (ii) Conc. HNO_3	(1) 
(Q)  (i) Conc. HNO_3/ Conc. H_2SO_4 (ii) Sn/HCl (iii) $NaNO_2$/HCl, 0 - 5°C, (iv) H_2O (v) Conc. HNO_3/ Conc. H_2SO_4	(2) 
(R)  (i) Conc. H_2SO_4 (ii) Conc. HNO_3 (iii) H_3O^+, Δ	(3) 
(S)  (i) (a) $KMnO_4$/ KOH, Δ; (b) H_3O^+ (ii) Conc. HNO_3/ Conc. H_2SO_4, Δ (iii) (a) $SOCl_2$, (b) NH_3 (iv) Br_2, NaOH (v) $NaNO_2$/HCl, 0 - 5° C, (vi) H_2O	(4) 
	(5) 

(A) P → 2; Q → 3; R → 4; S → 5 **(C)** P → 3; Q → 5; R → 4; S → 1
(B) P → 2; Q → 3; R → 5; S → 1 **(D)** P → 3; Q → 2; R → 5; S → 4

Answer Key

Q.No.	Answer key	Chapter's name	Topic's name
1	(D)	Root Mean Square	State of Matter
2	(A)	Disproportionation Reaction	p Block Element
3	(B)	Artificial Sweetner	Biomolecule
4	(C)	Isomers	Coordination Compound

5	(A,B,C)	Models	Atomic Strucutre
6	(A,B,D)	Chemical Reactions'	Phenols
7	(A,D)	Octet Rule	Chemical Compound and Strucutre
8	8120	V-T Graph	Thermodynamics
9	3	Order	Chemical Kinetics
10	18	Canizzaro Reaction	Aldehyde
11	3	Isoelectronic Species	Coordination Compound
12	909	Mix Reactions	Esters
13	1	Diamagnetic	Coordination Compound
14	(C)	Titrations	Electrochemistry
15	(B)	VSEpr	Chemical Bonding and Strucutre
16	(A)	Mix Reactions	Organic Chemistry
17	(C)	Mix Reactions	Organic Chemistry

❑❑

ANSWERS WITH EXPLANATIONS

1. Correct option is (D).

For Ideal Gas

PV = nRT

$\therefore n \propto$ **P at constant T & V.**

$\because$ **mole** $= \dfrac{\text{Mass}}{\text{Molar mass}}$

For gas X : $\dfrac{10}{M_X} \propto 2$ atm ...(1)

For gas X & Y : $\dfrac{10}{M_X} + \dfrac{80}{M_Y} \propto 6$ atm ...(2)

From (2) – (1)

$\dfrac{80}{M_Y} \propto 4$ ...(3)

On dividing (1) by (3)

$\dfrac{M_Y}{8M_X} = \dfrac{1}{2}$

$\therefore \dfrac{M_Y}{M_X} = 4$ ...(4)

$\because V_{rms} = \sqrt{\dfrac{3RT}{M}} \Rightarrow V_{rms} \propto \dfrac{1}{\sqrt{M}}$

$\therefore \dfrac{(V_{rms})_X}{(V_{rms})_Y} = \sqrt{\dfrac{M_Y}{M_X}} = \sqrt{\dfrac{4}{1}} = \dfrac{2}{1}$

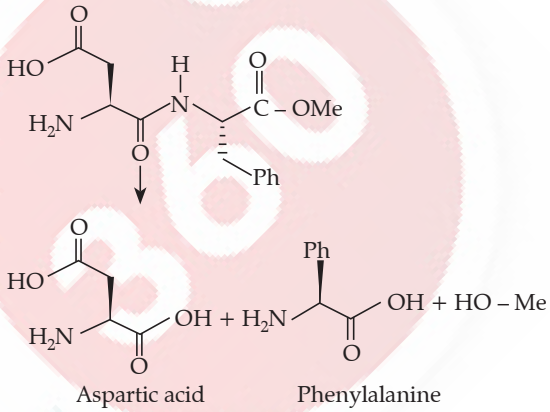
2. Correct option is (A).

$3HNO_2 \rightarrow HNO_3 + 2NO + H_2O$

$3HNO_2 \rightarrow NO_3^- + 2NO + H_3O^+$

3. Correct option is (B).

Aspartame structure is a dipeptide consisting aspartic acid and methyl ester of phenylalanine



Aspartic acid Phenylalanine

4. Correct option is (C).

Set-I: **$[Co(NH_3)_3(NO_2)_3]$ shows two geometrical isomers: facial and meridional**

$[Co(en)_2Cl_2]$ shows two geometrical isomers: cis and trans

Set-II: **$[Co(NH_3)_5Cl]SO_4$ and $[Co(NH_3)_5SO_4]Cl$ are ionisation isomers of each other.**

5. Correct option is (A, B, C).

(A) Exact path of electron in atom can't be defined with it's speed simultaneously only probability of finding electron is defined as orbitals.

(B) In 2s orbital energy of electron is –ve but at $r \rightarrow \infty$, energy (E) $\rightarrow$ 0

(C) As per Bohr's Model

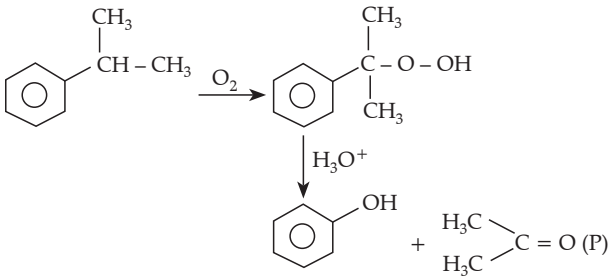
$E = -13.6\dfrac{z^2}{n^2}$ ev

at n = 1, E = –13.6 Z^2 ev (most negative)

(D) $V = 2.18 \times 10^6 \dfrac{z}{n}$ m/sec. **(by Bohr's model)**

$V \propto \dfrac{1}{n}$

6. Correct options are (A, B, D).



(A) $CHCl_3$ + $(H_3C)_2C=O$ $\xrightarrow{KOH}$ $(H_3C)_2C(OH)CCl_3$ (Chloritone)

(B) $CHCl_3$ (R) $\xrightarrow{O_2}$ $COCl_2 + HCl$ Phosgene gas

(C) Q does not undergo Cannizaro reaction

(D) $(CH_3-C(=O)-O)_2Ca \xrightarrow{\Delta} CH_3-C(=O)-CH_3 + CaCO_3$

7. Correct options are (A, D).

Compounds following octet rule are:

(A) CO_2, C_2H_4 & HCl

(B) O_3, HCl

(C) None

(D) CO_2, O_3, C_2H_4

8. Correct answer is [8120 J].

$\Delta H = nC_{p,m}\Delta T$

for xy process, $\Delta T = 0 = \Delta H_{xy} = 0$

for yz process, $\Delta T = 80$ K

$C_{p,m} = C_{v,m} + R = (12 + 8.3)$ J/k. mol.

$= 20.3$ J/mol.k

$\Delta H_{yz} = nC_{pm}\Delta T$

$= (5)(20.3)(80)$ J

$= 8120$ J

$\Delta H_{x\to y\to z} = \Delta H_{xy} + \Delta H_{yz}$

$0 + 8120$

$\therefore \Delta H_{x\to y\to z} = 8120$

9. Correct answer is [3].

Slow reacton is R.D.S and

Rate law of R.D.S = Rate law of reaction

$\Rightarrow$ **ror** $= k_2[N_2O_2][H_2]$ ...(1)

from fast equilibrium

$$\frac{k_1}{k_{-1}} = \frac{[N_2O_2]}{[NO]^2}$$

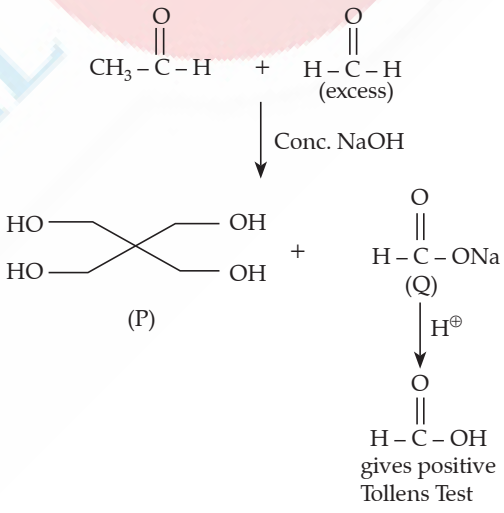
$$\Rightarrow [N_2O_2] = \frac{k_1}{k_{-1}}[NO]^2$$

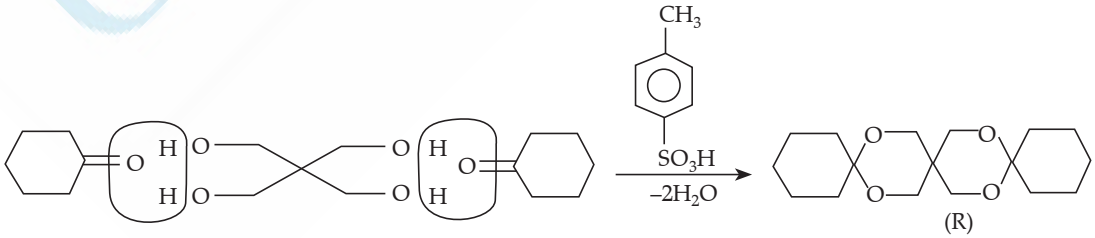
substituting in (1)

$$\Rightarrow \text{ror} = \frac{k_1k_2}{k_{-1}}[H_2][NO]^2$$

$\Rightarrow$ **Order of reaction is (3)**

10. Correct answer is [14 + 4 = 18].





Total CH_2 in R = 14 and oxygen in R = 4

So 14 + 4 = 18

11. Correct answer is [3].

Electrons in $Ni(CO)_4 = 28 + 4 \times 2 = 36$

Electrons in $Cr(CO)_5 = 24 + 2 \times 5 = 34$

Electrons in $Cu(CO)_3 = 29 + 2 \times 3 = 35$

Electrons in $Mn(CO)_5 = 25 + 2 \times 5 = 35$

Electrons in $Fe(CO)_5 = 26 + 2 \times 5 = 36$

Electrons in $[Co(CO)_3]^{4-} = 27 + 2 \times 3 + 3 = 36$

Electrons in $[Cr(CO)_4]^{4-} = 24 + 2 \times 4 + 4 = 36$

Electrons in Ir $(CO)_3 = 77 + 3 \times 2 = 83$

So iso-electronic species which are similar to $Ni(CO)_4$ are $Fe(CO)_5$, $[Co(CO)_3]^{4-}$, $[Cr(CO)_4]^{4-}$

So answer is 3.

12. Correct answer is [909].

$$H-C\equiv C-(CH_2)_{15}-CO_2Et$$
$$\downarrow Hg^{2+}/H_3O^+$$
$$H_3C-\overset{\overset{O}{\|}}{C}-(CH_2)_{15}-CO_2Et$$
$$\downarrow Zn(Hg) + \text{Conc. HCl}$$
$$H_3C-CH_2-(CH_2)_{15}-CO_2Et$$
$$\downarrow H_3O^+$$
$$H_3C-(CH_2)_{16}-CO_2H$$
$$\downarrow \text{Glycerol}/H^+$$

Glyceryl stearate

$$\begin{array}{l} CH_2-O-\overset{\overset{O}{\|}}{C}-C_{17}H_{35} \\ | \\ CH-O-\overset{\overset{O}{\|}}{C}-C_{17}H_{35} \\ | \\ CH_2-O-\overset{\overset{O}{\|}}{C}-C_{17}H_{35} \end{array} \quad (Q)$$

(1 mole)

$$\downarrow NaOH(\text{excess})$$

$$\begin{array}{l} CH_2OH \\ | \\ CHOH \\ | \\ CH_2OH \end{array} + 3C_{17}H_{35}CO_2Na$$

(3 mole)

$$\downarrow CaCl_2$$

$$(C_{17}H_{35}CO_2)_2Ca \text{ (Soap)}$$

3/2 mole

The amount of *R* produced in gram

$$= \frac{3}{2} \times [34 \times 12 + 70 \times 1 + 2 \times 44 + 40]$$

$$= \frac{3}{2} \times 606$$

$$= 909$$

13. Correct answer is [1].

$[Mn(NH_3)_6]^{+3} \Rightarrow 3d^4$ + S.F.L $\Rightarrow$ two unpaired electrons

$[MnCl_6]^{3-} \Rightarrow 3d^4$ + W.F.L $\Rightarrow$ four unpaired electrons

$[FeF_6]^{3-} \Rightarrow 3d^5$ + W.F.L $\Rightarrow$ five unpaired electrons

$[CoF_6]^{3-} \Rightarrow 3d^6$ + W.F.L $\Rightarrow$ four unpaired electrons

$[Fe(NH_3)_6]^{+3} \Rightarrow 3d^5$ + S.F.L $\Rightarrow$ one unpaired electron

$[Co(en)_3]^{+3} \Rightarrow 3d^6$ + S.F.L $\Rightarrow$ zero unpaired electron & diamagnetic.

14. Correct option is (C).

Option (P) : On adding $AgNO_3$ solution to KCl solution precipitation of AgCl will occur due to which Cl^- already present will be replaced by NO_3^- ions. So conductance of solution will decrease till equivalence point. After complete precipitation of AgCl, further added $AgNO_3$ will increase the number of ions in resulting solution so conductance will increase.

Option (Q) : On adding KCl solution to $AgNO_3$ solution precipitation of AgCl will occur due to which already present Ag^+ ions will be replaced by K^+ ions in solution. So, conductance of solution will increase. After complete precipitation of AgCl further added KCl will increase the number of ions in resulting solution so conductance will increase further.

Option (R) : On adding HCl solution to NaOH solution, OH^- will be replaced by Cl^- ions so conductance of solution decreases. After complete neutralisation further added HCl will increase number of ions in the solution. So, conductance will increase further.

Option (S) : On adding CH_3COOH solution to NaOH solution OH^- will be replaced by CH_3COO^- ions, so conductance of solution decreases. After complete neutralisation further added CH_3COOH will remain undissociated because it is a weak acid and there is also common ion effect on acetate ions. So number of ions in solution will remain almost constant therefore conductance of solution will remain constant.

15. Correct option is (B).

$XeF_2 \Rightarrow$ 2 sigma bonds and 3 lone pairs on Xe, number of hybrid orbitals = 5, sp^3d hybridisation , geometry will be trigonal bipyramidal.

P–5

$XeF_4 \Rightarrow$ 4 sigma bonds and 2 lone pairs on Xe, number of hybrid orbitals = 6, sp^3d^2 hybridisation , geometry will be octahedral.

Q–3

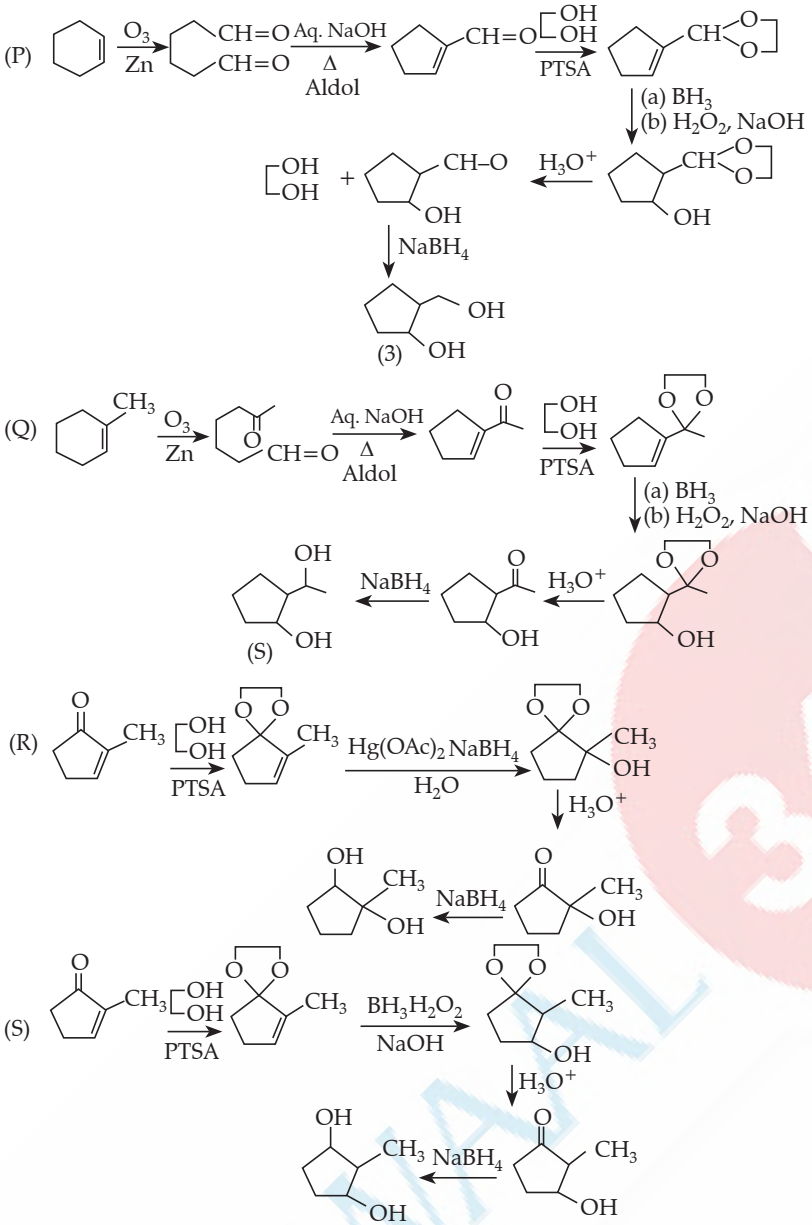
$XeO_3 \Rightarrow$ 3 sigma bonds and 1 lone pairs on Xe, number of hybrid orbitals = 4, sp^3 hybridisation , geometry will be tetrahedral.

R–2

$XeO_3F_2 \Rightarrow$ 5 sigma bonds and 0 lone pairs on Xe, number of hybrid orbitals = 5, sp^3d hybridisation , geometry will be trigonal bipyramidal.

S–4

16. Correct option is (A).



17. Correct option is (C).